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13. All acid—base reactions have two conjugate acid—base pairs. One conjugate acid—base pair in the reaction in Model 3 is $\text{H}_2\text{O}/\text{H}_3\text{O}^+$. List the other acid—base pair in the reaction. 14. Why is HCO_3^- considered the

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"acid" part of the pair in the reaction in Model 3? 15. is CO considered the "base" part of the pair in the reaction in Model 3? 16.

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Acids and Bases Model
1: Two definitions of
Acids and Bases
Arrhenius Definitions
An acid is a substance
that produces
hydronium ions,
 $\text{H}_3\text{O}^+(\text{aq})$, when it is

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added to water. A base is a substance that produces hydroxide ions, $\text{OH}^-(\text{aq})$ when it is added to water.

Brønsted-Lowry

Definitions An acid is a substance that donates a proton.

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Acids and bases react with one another to yield two products: water, and an ionic compound known as a salt. This kind of

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reaction is called a neutralization reaction.

10.1: Introduction to
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We can't discuss acids
and bases without

talking about pH. pH
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Model 3 — Conjugate
Acid—Base Pairs Base
+ 1-120 (l) HC031-(aq)

13. All acid—base
reactions have two
conjugate acid—base
pairs. One conjugate
acid—base pair in the
reaction in Model 3 is $\text{H}^+/\text{H}_2\text{O}$. List the other
acid—base pair in the
reaction. 14. Why is
FICO - considered the
"acid" part of the pair

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in the reaction in Model
3? 15.

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and Base Equilibria Key
Questions 1. A 0.0100
M solution of a weak
acid HA has a pH of

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2.60. What is the value of K_a for the acid?
[Hint: What is the actual concentration of undissociated HA, $[HA]$, in this solution?] $pH = 2.60$ $[H_3O^+] = 2.5 \times 10^{-3} M$

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10 - Solutions Weak
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