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Flashcards. Learn. Write. Spell. Test.
PLAY. Match. Gravity. Created by.
adriannamnelson15. Terms in this set
(19) Ion. An atom that gained or lost
electrons and has a charge. Cation-Lost
electron-Positive charge-Metals. Anion-
Gained electrons-Negative charge-

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Nonmetals. Properties.

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The energy required to separate 1 mol of ions in an ionic compound is referred to as lattice energy. Lattice energy is directly related to the size of the ions that are bonded. Smaller ions form

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compounds with more closely spaced ionic charges and require more energy to separate. Smaller, harder to get between them.

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206 Chapter 7 • Ionic Compounds and Metals Section 7.1.1 Figure 7.1 As

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carbon dioxide dis-solves in ocean water, carbonate ions are produced. Coral polyps capture these carbonate ions, producing crystals of calcium carbonate, which they secrete as an exoskeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

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Chapter 7: Ionic Compounds and Metals - Mr. Miller

Chapter 7 Review: Ionic Compounds and Metals. Place the response in the blank, at the left, that best completes the following: _____1. _____ is the force that holds 2, or more, atoms together. _____2. _____ electrons are those electrons that are in the outer energy level. _____3.

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level of an.... In forming compounds,

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confi....

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SOLUTIONS MANUAL 25. calcium and

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chlorate $\text{Ca}(\text{ClO}_3)_2$ 26. aluminum and carbonate $\text{Al}_2(\text{CO}_3)_3$ 27. Challenge

Write the formula for an ionic compound formed by ions from a group 2 element and polyatomic ions composed of only of carbon and oxygen. The polyatomic ion is carbonate (CO_3^{2-}).

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Compounds and Metals

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COMPOUNDS 203 SECTION 7-1 O

OBJECTIVES Explain the significance of a chemical formula. Determine the formula of an ionic compound formed between two given ions. Name an ionic compound given its formula. Using prefixes, name a binary molecular compound from its formula. Write the formula of a binary molecular compound

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given its name. C

CHAPTER 7 Chemical Formulas and Chemical Compounds

A compound that contains ions and is held together by ionic bonds is called an ionic compound. The periodic table can help us recognize many of the compounds that are ionic: When a metal

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is combined with one or more nonmetals, the compound is usually ionic.

3.7 Molecular and Ionic Compounds - General Chemistry 1 & 2

A compound that contains ions and is held together by ionic bonds is called an ionic compound. The periodic table can

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help us recognize many of the compounds that are ionic: When a metal is combined with one or more nonmetals, the compound is usually ionic.

2.7 Molecular and Ionic Compounds | Chemistry

Naming Ions. The name of a monatomic

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cation is simply the name of the element followed by the word ion. Thus, Na^+ is the sodium ion, Al^{3+} is the aluminum ion, Ca^{2+} is the calcium ion, and so forth.. We have seen that some elements lose different numbers of electrons, producing ions of different charges (Figure 3.3).

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**5.7: Naming Ionic Compounds -
Chemistry LibreTexts**

Ionic Bonds - Chapter 7 1. Ionic Bonding and Ionic Compounds Or How I Learned to Love Electrostatic Forces 2. Ionic Bonds The periodic table is based upon the properties of the elements. After the discovery of the electron, scientists found that all elements in the same

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group have the same number of valence electrons.

Ionic Bonds - Chapter 7

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Chapter 7 Ionic and Metallic Bonding 157
25. State the number of electrons lost or

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gained in forming each of these ions.
Name the ions and tell whether it is an
anion or a cation.

**05 CTR ch07 7/9/04 3:27 PM Page
155 IONS 7**

Yay, this practice test will determine if
you're ready to take the real thing.
Questions are taken right out of the

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book/packet and review lotsa vocab.

Chapter 7 Chem: Practice Test - ProProfs Quiz

ions attract each other, forming ionic bonds. The resulting compounds are called ionic compounds and are the primary subject of this chapter. The second way for an atom to obtain an

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octet of electrons is by sharing electrons with another atom. These shared electrons simultaneously occupy the outermost shell of more than one atom. The bond

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