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Solid-State Chemistry Class 12 NCERT Solutions: Question 3 and 4. Through these two answers, students will be able to classify solid as ionic, metallic, molecular, covalent or amorphous. Students will also learn the meaning of coordination number in both cubic close-packed structure and body centred cubic structure.

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www.ncerthelp.com (Visit for all ncert solutions in text and videos, CBSE syllabus, note and many more) Chemistry Notes for class 12 Chapter 1 The Solid State Solids Solids are the chemical substances which are characterised by define shape and volume, rigidity, high density, low compressibility. The constituent particles (atoms, molecules or ions)

### Chemistry Notes for class 12 Chapter 1 The Solid State

Question: 1.7 - Solid A is a very hard electrical insulator in solid as well as in molten state and melts at extremely high temperature. What type of solid is it? Answer: Given solid 'A' is a covalent solids, such as diamond. Question: 1.8 - Ionic solids conduct electricity in molten state but not in solid state. Explain.

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Chemistry 2 In Class XI you have learnt that matter can exist in three states namely, solid, liquid and gas. Under a given set of conditions of temperature and pressure, which of these would be the most stable state of a given substance depends upon the net effect of two opposing factors.

### Objectives The Solid State - Prashanth Ellina

Solid State NCERT Solution (part 1) Question – 1.1 - Define the term 'amorphous'. Give a few examples of amorphous solids. Answer: Solids having constituent particles with irregular shapes and short range order are called amorphous solids. Amorphous solids are isotropic in nature and melt over a range of temperature.

### 12 Chemistry Solid State NCERT Exercise Solution part 1

CBSE Class 12 Chemistry Quick Revision Notes Chapter 1 The Solid State. Solid: Solid is a state of matter in which the constituting particles are arranged very closely. The constituent particles can be atoms, molecules or ions. Properties of solids: They have definite mass, volume and shape. They are compressible and rigid.

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Solid state chemistry class 12 ncert solutions. At sufficiently low temperatures the thermal energy is low and the intermolecular forces close them so much that they attach to each other and the definite positions are absorbed. It can still oscillate around its mean positions and the substance remains in solid-state.

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