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in your school would

have provided an

excellent foundation in

functional group

chemistry. It is

therefore not the

intention of the present

module (PR1110 -

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cover in detail the

theory of functional

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group chemistry once  
again.

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chemistry.

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by Charles Dickson**

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IIA INTRODUCTION The

practical part of the  
course covers

experiments on the

synthesis of organic

compounds and drugs

using different analysis

and purification

techniques and

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includes synthesis of chiral drugs and the techniques used in the separation of enantiomers.

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Analytical Balances

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of Savitribai Phule Pune  
university semester V.

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Dickson Step 1: First

step involves the

formation of Acetylium

ion, the reaction is

enhanced by proton

donation by

concentrated acid.

Step 2: Amino group of

Para-Aminophenol

dissociates as follows:

Step

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i\_lab\_manual 1.

EXPERIMENT NO:

MEDICINAL CHEMISTRY

I BNPCW SYNTHESIS OF

BARBITURIC ACID

FROM DIETHYL

MALONATE AIM: To

prepare and submit

barbituric acid from

diethyl malonate.

APPARATUS: Round

bottom flask, Reflux

condenser, Beaker,

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Glass rod, Buchner  
funnel, Filtration  
apparatus.

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Pharmacognosy

&Phytochemistry Lab.

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academic calendar.

Week Basic and

support material to be  
covered. 16-20/10

Safety rules 23-27/10

Introduction to the

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Microscopy

30/10-3/11

Microscopical

identification for starch  
+ calcium oxalate

6-10/11 Microscopical

identification for ginger  
root + cinnamon barks

13-17/11 Microscopical  
identification for senna

leaves + chamomile  
flower.

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& Phytochemistry  
Lab.**

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## Practical Inorganic II Chemistry (PDF 69P)

This note describes the laboratory activities to understand the model of ionic compound close packing, reduction-oxidation reaction of several metals, improve the double salt preparation and re-crystallization skills as the basic for advance laboratory activities.

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following laboratories

have been compiled

and adapted by Alan

Khuu, M.S. & Armando

Rivera, Ph.D.

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Step 1: First step involves the formation of Acetylium ion, the reaction is enhanced by proton donation by concentrated acid.

Step 2: Amino group of Para-Aminophenol dissociates as follows:

Step 3: Acetylium ion

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attacks the anion at  
the nitrogen atom to  
form Para-acetyl  
Aminophenol or  
Paracetamol.

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Medicinal Chemistry-  
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laboratory procedures.

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exercises focusing on  
common automated  
and manual clinical  
chemistry testing  
procedures for glucose,  
electrolytes,

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practical organic  
chemistry, we have  
extended the scope of  
the work so that it

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covers most of the needs of students working for an Honours or Special Degree. To meet the needs of the advanced students, preparations have now been included to illustrate, for example, reduction by lithium aluminium hydride and by the Meerwein-Ponndorf-

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## **Of Practical Inorganic Chemistry**

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In a 250 mL volumetric flask, prepare 0.1 M citric acid solution.

Load a burette with 50 mL of the 0.2 M sodium phosphate dibasic solution. Set up the burette on a retort stand with a burette clamp. Measure out 100 mL of citric acid using a 100 mL graduated cylinder.

Pour into a 250 mL PHC

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